

Chemical and isotopic composition of fumarolic gases at Iwate volcano, Japan, during and after seismic activity in 1998: implications for the modification of ascending volcanic gases

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ABSTRACT

In 1998, there were many volcanic earthquakes recorded at Iwate volcano, Japan. Although an eruption was anticipated, it never occurred. Fumarolic gases were sampled at the volcano on six occasions during 1998 and 1999, and were analyzed for their chemical compositions and isotope ratios. The fumarolic gases were again sampled in 2004 and 2006, after the period of seismic activity. The HCl concentrations and isotope ratios of the H₂O in the gas samples collected from an active geothermal area at the volcano were high in 1998 and 1999, but decreased significantly after 2004, irrespective of the relatively stable concentrations of CO₂ and sulfur-bearing gases. A notable feature of the fumarolic gases is the high isotope ratio of H₂O, similar to that of pure magmatic gas, which is typical of andesitic volcanoes, although the temperature at the outlet was less than 146 °C in the present case. These features can be explained if the enthalpy of the magmatic gas was equivalent to that of water vapor at 252 °C to 370 °C under a pressure of 1 bar. The correlations observed among CO₂/H₂O and HCl/H₂O, and the δ¹⁸O values of the fumarolic gases suggest partial condensation of H₂O vapor during the ascent of the volcanic gas to the surface. The estimated CO₂/H₂O molar ratio of the magmatic gas was 0.008, which is less than the 0.03 reported previously for magmatic gas sampled during an effusive eruption at Unzen volcano, Japan. These data suggest that the magma at Iwate volcano is depleted in volatiles, because CO₂ is preferentially degassed from the silicate melt relative to H₂O. Such depletion in CO₂ might explain the failed eruption at Iwate volcano in 1998.

1. Introduction

Fumarolic gases generally include volatiles that are emitted from degassing magma; consequently, the chemical and isotopic compositions of such gases can provide information on the degassing magma. However, the chemical and isotopic compositions of the gases can be modified during

the ascent to the surface, due to, for example, contamination by ground water, condensation of water vapor, and removal of acidic components, such as HCl.

In 1998, many volcanic earthquakes occurred at Iwate volcano, Japan. However, although an eruption was anticipated, it never occurred. During and after this period of volcanic seismicity, we sampled fumarolic gases at the volcano and measured their chemical and isotopic compositions. In this study, we investigate the processes that acted to modify the chemical compositions and isotope ratios of these fumarolic gases, and we estimate the CO₂/H₂O ratios of the magmatic gases. The modification process is quantitatively modeled assuming simple physicochemical processes, which successfully reproduce the observed variations in the chemical and isotopic compositions of these fumarolic gases.

1.1. Iwate volcano

Iwate volcano is an active volcano in northern Honshu, Japan (Figure 1). The volcano has two topographic peaks and it is divided into its western and eastern parts, which are centered around Onawashiro crater (Figure 1, O) and Yakushidake crater (Figure 1, Y), respectively. Itoh and Doi [2005] investigated the development of Iwate volcano. The eastern part is younger than the western part, and its surface layer was deposited over the past 7,000 years. An eruption occurred at the outer rim of the Yakushidake crater in 1732, which produced lava flow. The eastern part contains limited geothermal activity. A weak and diffusive steam discharge occurred at point H in Figure 1, without any apparent sulfur-bearing gases. The western part is relatively old, with the most recent lava from the Onawashiro crater estimated to be older than 30 ka [Itoh and Doi 2005]. The western part is characterized by intensive geothermal activity that is

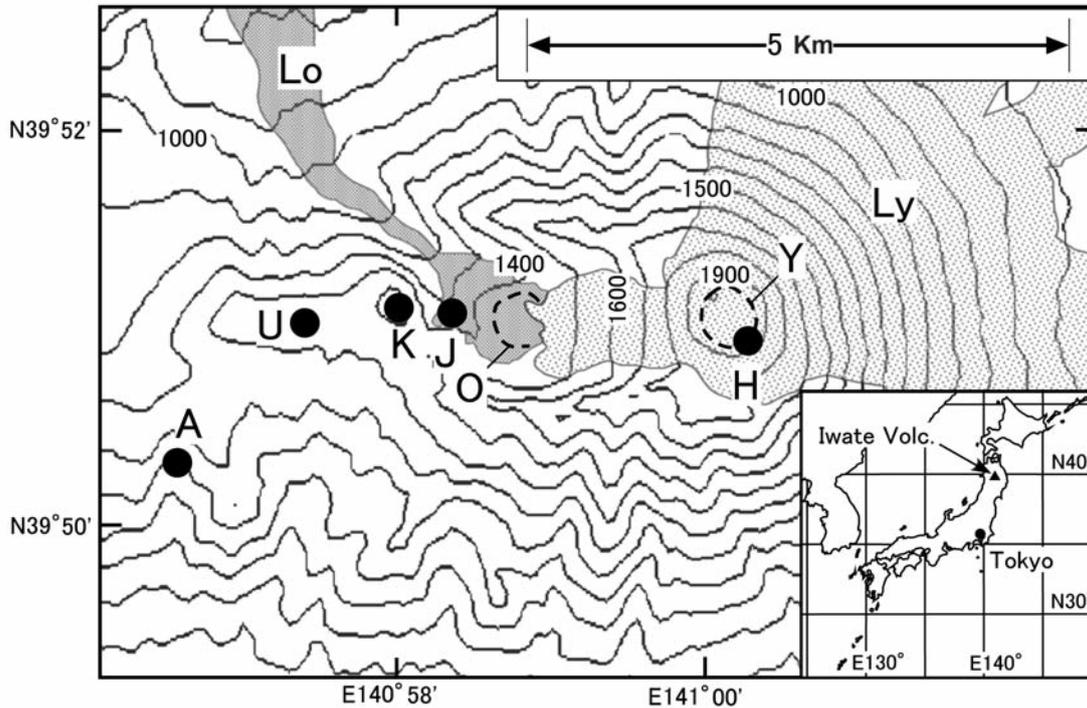


Figure 1. Topographic map of Iwate volcano (contour interval, 100 m; altitude, meters above sea level). Inset, location of the volcano. Dashed curves O and Y, outer rims of the Onawashiro and Yakushidake craters, respectively. Shaded areas Lo and Ly, distribution of the lava flows from the Onawashiro and Yakushidake craters, respectively. Letters H (Higashi iwate), J (Ojigokudani), K (Kurokura), U (Ubakura), and A (Amihari), locations of geothermal areas from which the fumarolic gases were collected. At site J ($39^{\circ}51'05''\text{N}$, $140^{\circ}58'21''\text{E}$), hot spring water was also collected. The area with altered soil at site J measured about 100 m (north–south) by 75 m (east–west). Intense fumarolic activity was noted at J, including a strong smell of sulfur-bearing gases (also detected at site A). At sites H, K and U, in contrast, sulfur-bearing gases were not smelt when sampling the fumarolic gases.

concentrated in four areas (Figure 1, J, K, U and A). At site J, in particular, there was a strong smell of sulfur-bearing gas. As the most recent volcanic activity, a steam driven explosive eruption was observed at site J in 1919.

About 300,000 people live in Morioka city, which is located 20 km southeast of Iwate volcano, and a north–south express highway is located 9 km east of Iwate volcano. During an eruption at Yakushidake crater in 1686, houses near Iwate volcano were affected by ash deposits and a lahar. Any future eruption at Iwate volcano poses serious risk to nearby residents and infrastructure.

Iwate volcano is monitored continuously by the Japan Meteorological Agency, in terms of its seismicity and crustal deformation (tiltmeter and GPS), but its fumarolic gases are not monitored. In terms of the geochemistry of the fumarolic gases at Iwate volcano, only Noguchi et al. [1961] have reported on the chemical composition of the gas and sublimate within the Yakushidake crater.

Miura et al. [2000] reported deformation of the volcanic body at Iwate volcano in 1998, which was attributed to the intrusion of magma at a depth of 3 km beneath the western part. From May to September 1998, many volcanic earthquakes were recorded beneath the region covered by points O, J, K and U in Figure 1 [Miura et al. 2000]. The earthquake hypocenters were concentrated in a zone that trended E–W.

Tanaka et al. [2002] reported a region of anomalously high seismic velocity beneath the Onawashiro crater (Figure 1), which was attributed to a body of solidified magma. A similar model was proposed by Aizawa et al. [2009] based on a magnetotelluric survey.

2. Sampling and analysis of gas and water

Most of the fumarolic gases at sites H, J, K, U and A (Figure 1) were sampled following the method proposed by Ozawa [1968], using a glass syringe in which 20 ml of 5 M KOH solution was filled, to absorb acidic gases. This method results in only a small amount of air contamination in the syringe. The KOH solution obtained was subjected to wet chemical analysis, according to Ozawa [1968]. The gas phase collected in the syringe, which included the «residual» composition of the gas (e.g., N_2 , O_2 , H_2 , CH_4 , Ar, He and Ne), was analyzed using two gas chromatographs (GCs) with Ar and O_2 as the carrier gases, through which the components other than Ar and O_2 were analyzed. The GC data were then combined to yield the complete composition of the residual gas. Fumarolic gases at site J were sampled in 2004 and 2006 for He isotope analysis, using an evacuated glass bottle [Giggenbach and Goguel 1989] that included 20 ml 5 M KOH, to prevent air contamination. At site J, hot spring water was also sampled in 1998 and 1999. Two fumaroles at site J were selected for repeated sampling. One series of samples (Table 1,

FUMAROLIC GASES AT IWATE VOLCANO, JAPAN

Location	Code	Date	Temp °C	H ₂ O	CO ₂	H ₂ S	SO ₂	HCl	He	H ₂	O ₂	N ₂	CH ₄	Ar	δD _{S_MO_W} ‰	δ ¹⁸ O _{S_MO_W} ‰	δ ¹³ C _{P_DB}	³ He/ ⁴ He R/Ra	⁴ He/ ²⁰ Ne
J	JA1	1998/6/2	135	980200	13600	5240	268	958	0.032	94	**	107	0.32	0.29	-31	3.2	-3.2		
J	JB1	1998/6/2	110	984100	11900	3210	166	612	0.028	1.2	**	82.7	0.081	0.254	-27	3.1	-3.5		
J	JA2	1998/7/24	134	986500	9280	3410	357	1170	0.015	2.3	**	45.8	0.29	0.17	-26	4.4	-3.3		
J	JB2	1998/7/24	106	986900	9540	2350	279	926	0.019	0.41	**	52.8	0.052	0.21	-25	4.4	-3.2		
J	JA3	1998/9/10	135.6	984900	9620	3800	545	2280	0.014	2.5	**	52.2	0.025	0.18	-22	5.4	-3.0		
J	JB3	1998/9/10	109	986200	10500	2610	465	447	0.023	2.4	**	67.8	0.057	0.26	-25	4.4	-3.0		
J	JA4	1999/5/12	143	983300	8420	3150	743	3990	*	345	**	89.0	0.37	*	-22	5.2	-3.5		
J	JB4	1999/5/12	102.6	987900	9170	2000	520	390	0.015	0.61	**	51.6	0.055	0.34	-25	4.8	-3.2		
J	JA5	1999/8/17	139	981700	10110	3570	966	3490	*	67.2	**	62.3	0.17	*	-24	4.8			
J	JB5	1999/8/17	98	988500	8800	1930	406	298	0.0065	0.27	**	25.1	0.023	0.14	-22	4.8			
J	JA6	1999/10/14	146	981500	11000	5000	1240	1110	*	72.3	**	79.7	0.24	0.43	-30	3.4			
J	JB6	1999/10/14	101	986300	10400	2730	552	ud	0.023	1.03	**	64.9	0.11	0.23	-29	3.5			
J	JA7	2004/9/22	111.3	979100	15600	4110	508	663	0.0059	1.1	1.8	35.3	0.25	0.16	-34	0.0	-3.9	5.01	46.4
J	JB7	2004/9/22	96.3	967200	28900	3720	91	ud	0.0066	0.8	1.4	40.1	0.29	0.22	-46	-2.7	-3.3	4.91	313
J	JA8	2006/8/1	116	254100	14200	5830	1730	ud	ud	4.0	127000	590000	ud	6800	-53	-3.5			
J	JB8	2006/8/1	96.6	984100	11200	4460	105	ud	0.021	1.6	ud	74.7	0.00012	1.09	-81	-10.3	-2.7	4.82	31.6
U	U1	1998/9/10	94.6	870700	129300										-148	-26.0	-3.5		
U	U2	2006/8/2	95.6	475100	13300					3.9	101000	406000	ud	4660	-135	-25.5	-4.9	1.08	0.34
K	K	2006/8/2	92.4	845100	4400					ud	29500	120000	ud	1390	-121	-20.8	-4.2	1.12	0.33
A	A	2005/9/5	121.3	988100	7850	3670	28.0	ud	*	95	21	206	50.7	1.70	-61	-4.3	-4.4	4.32	82.9
H	H1	2005/9/1	83.8	312000	6680					3.8	137000	538000	ud	6270	-112	-17.1	-7.6	1.04	0.34
H	H2	2005/9/1	83.6	491300	13900					9.8	97800	392000	ud	4590	-139	-22.0	-5.9	1.11	0.33
Y	Y1	1960/9/29	144	983000	14600	1090	817	376											
Y	Y2	1960/9/29	183	982000	18400	1170	1520	185											

Table 1. Chemical and isotopic compositions of the fumarolic gases at Iwate volcano.

Chemical concentrations as μmol/mol. ud, below detection limit. Until October 1999, the gases were sampled using a glass syringe as this method was easy to apply and sufficiently accurate in terms of the major chemical components, except He and Ar. After 2004, the gases were sampled using an evacuated glass bottle to measure the He isotope ratio. *He concentration not measured due to interference from a large peak of H₂ in the gas chromatograph. **O₂ subtracted from the raw data assuming the O₂ detected was contaminated by air trapped in the glass syringe used for sampling. Subtraction of the atmospheric component was performed for N₂ and Ar, based on the concentration of O₂. Empty cells, no analysis. R, ³He/⁴He molar ratio. Ra, R value of normal air as 1.4 × 10⁻⁶. The chemical compositions of the gases within the Yakushidake crater (Y1, Y2) are from Noguchi et al. [1961].

JA1–JA8) was collected at the fumarole with the highest outlet temperature. A second series (Table 1, JB1–JB8) was sampled at a fumarole located 32 m from the first. In addition, hot spring waters were sampled at site J for analysis of their anionic compositions by ion chromatography.

To measure the δD values of H_2O in the fumarolic gases, we sampled the gas condensates. Part of the condensates was converted to H_2 gas with metallic Zn [Coleman et al. 1982] for analysis using mass spectrometry (Finnigan MAT 252). The same procedure was used for measuring the δD value of H_2O in hot spring water. The $\delta^{18}O$ values of H_2O in the condensate and water sample were measured using the conventional CO_2 - H_2O equilibration method [Epstein and Mayeda 1953], using mass spectrometry. To measure the $\delta^{13}C$ value of CO_2 in the fumarolic gases, the CO_2 in the gases was fixed as a $BaCO_3$ precipitate at the sampling site [Kusakabe 2001]. The precipitate was converted to CO_2 in the laboratory for analysis by mass spectrometry. The analytical errors were 1.0‰, 0.1‰, and 0.1‰ for δD , $\delta^{18}O$ and $\delta^{13}C$, respectively. The gas collected in the headspace of the evacuated bottle was analyzed noble mass spectrometer (MM5400), to obtain the rare gas composition and $^3He/^4He$ ratio. The analytical error of the $^3He/^4He$ measurements was less than ± 0.1 Ra.

Analyses of chemical and stable isotopes were performed at the Kusatsu–Shirane Volcano Observatory at Gunma Prefecture, Japan, and those of He isotopes were performed at the Geological Survey of Japan at Tsukuba city, Japan.

3. Results

3.1. Major components of fumarolic gases

Tables 1 and 2 list the chemical and isotopic data for the fumarolic gases and hot spring waters, respectively. The major components of the fumarolic gases are H_2O , CO_2 , H_2S , SO_2 and HCl. Figure 2 shows a ternary diagram of the H_2O , CO_2 and S concentrations in the fumarolic gases (where $S = H_2S + SO_2$). Except for the samples JB7 and JA8, the gases at site J are clustered in the center of the diagram, along with the sole sample from site A. The gases at sites U, K and H are depleted in S; as such, they lie along the right-hand edge of the ternary diagram in Figure 2. The CO_2/H_2O ratios of samples from sites H and U are distinctly (although only slightly) higher than those of the cluster (except for sample U1). Sample U1 is strongly depleted in H_2O , whereas the sample from site H is enriched in H_2O . Gases at Yakushidake crater, as sampled in 1960, were enriched in S, similar to the composition of gases at site J.

3.2. Temporal variations in the compositions of the fumarolic gases at site J

Figure 3 shows the temporal variations in the chemical compositions and isotope ratios of the gases from site J.

The CO_2 and H_2S concentrations were relatively low in 1999, and they showed slight increases in 2004 and 2006. The SO_2 concentrations in the JA samples increased from June 1988 to October 1999, followed by a decrease in 2004 and an increase in 2006. Except for 2006, the temporal variations observed for the JB samples are similar to those for the JA samples. The HCl concentrations in the JA samples show an increase from June 1998 to May 1999, followed by a decrease. The HCl concentrations in the JB samples were relatively high in July 1998, but decreased gradually thereafter. The H_2 concentrations in the JA samples were high in May 1999, moderate in late 1999, and low in 2004 and 2006. The CH_4 concentrations in the JA and JB samples show no clear temporal trends. There is no correlation between the values of the JA and JB series. The δD and $\delta^{18}O$ values were high until August 1999, but decreased thereafter.

Code	Date	Temp °C	pH	δD_{SMOW} ‰	$\delta^{18}O_{SMOW}$ ‰	Cl^- mg/liter	SO_4^{2-} mg/liter
W1	1998/7/24	75.7	2.4	-69	-10.4	290	825
W2	1998/7/24	30.6	2.5	-73	-11.5	26	1820
W3	1998/7/24	70.1	1.0	-21	4.9	9300	4050
W4	1998/7/24	95.3	2.1	-64	-8.9	711	2520
W5	1998/9/10	94.8	1.2	-31	0.2	4309	2230
W6	1998/9/10	78.2	2.3	-68	-9.8	472	800
W7	1998/9/10	96.7	2.0	-53	-6.2	489	2211
W8	1998/9/10	90.9	2.1	-64	-9.3	442	730
W9	1999/5/12	74.2	1.9	-73	-10.9	14	1700
W10	1999/5/12	94.4	2.0	-60	-7.5	160	997
W11	1999/5/12	95.1	0.9	-18	4.8	8330	4790

Table 2. Chemical and isotopic compositions of the hot spring waters sampled at site J.

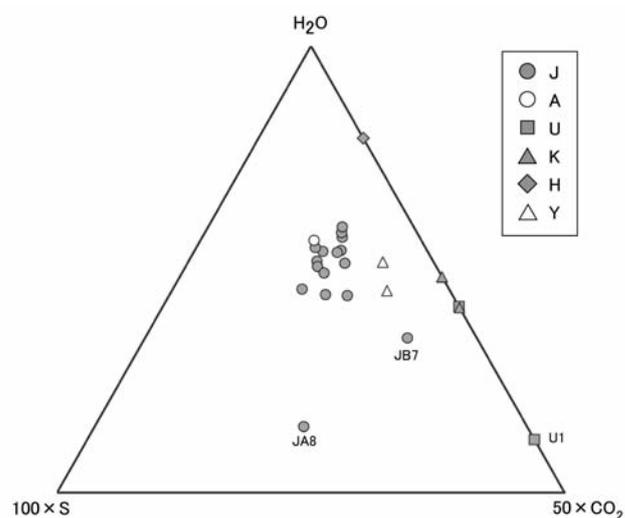


Figure 2. H_2O - CO_2 -S ternary diagram ($S = H_2S + SO_2$) showing the compositions of the fumarolic gases (molar basis).

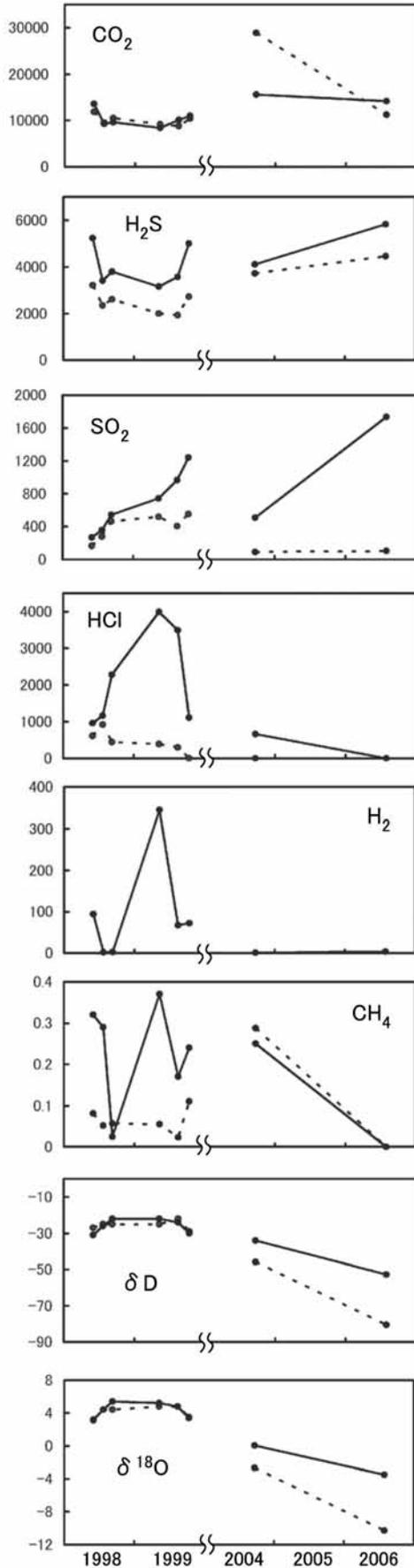


Figure 3. Temporal variations in the compositions ($\mu\text{mol/mol}$) and isotope ratios (‰ relative to SMOW) of the water in the fumarolic gases from site J (as indicated). Solid and dotted lines, JA and JB series, respectively.

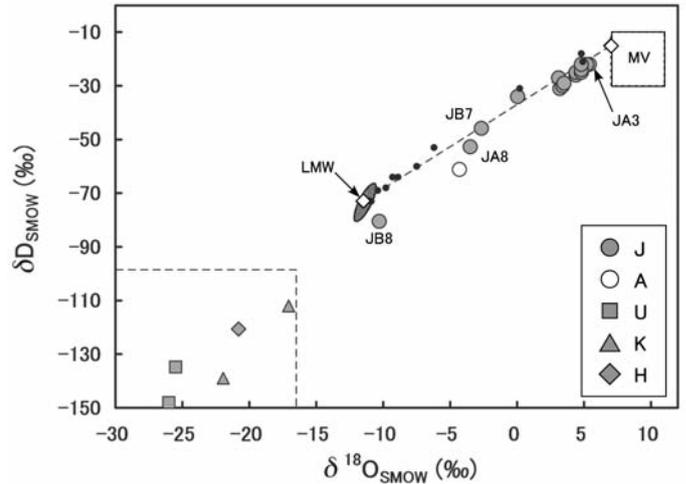


Figure 4. The δD and $\delta^{18}\text{O}$ values of water in the fumarolic gases and hot spring waters. Square, upper right, general range of the values for the typical MV of an andesitic volcano above a subduction zone. Elliptical region, local meteoric water (LMW) at the summit area of Iwate volcano [Kazahaya et al. 2000]. Small filled dots, isotope ratios of the hot spring water. Dashed horizontal and vertical lines, bottom left, δD and $\delta^{18}\text{O}$, respectively, of the water vapor equilibrated with the LMW at $100\text{ }^\circ\text{C}$.

3.3. Isotope ratios of H_2O in the gases and water

In 1999, the δD and $\delta^{18}\text{O}$ values of H_2O in the gases at site J were high, similar to those of magmatic water (typical of andesitic volcanoes) [Taran et al. 1989], as shown in Figure 4. The gases with high isotope ratios contained high concentrations of HCl (e.g., JA4). In Figure 4, the gases from site J are distributed along the line connecting the general magmatic water (Figure 4, square at top right) and local meteoric water, as reported by Kazahaya et al. [2000], who analyzed seven water samples from altitudes of 1,300 m to 1,800 m on the Iwate volcano. The average δD and $\delta^{18}\text{O}$ values of the seven samples were -73‰ and -11.5‰ , respectively.

Some of the hot spring waters (Table 2, W3 and W11) sampled at site J yielded high isotope ratios, close to those of magmatic water. The concentration of Cl^- in samples W3 and W11 was 9,300 mg/L and 8,330 mg/L, respectively, which is equivalent to 4,720 mmol/mol and 4,230 mmol/mol as the $\text{Cl}/\text{H}_2\text{O}$ molar ratio (similar to the $\text{HCl}/\text{H}_2\text{O}$ molar ratio for sample JA4: 4,060 mmol/mol). In Figure 4, the isotope ratios of the hot spring waters are also distributed along the line connecting the magmatic vapor (MV) and the local meteoric water (LMW).

The δD and $\delta^{18}\text{O}$ values of H_2O in the gases from sites U, K and H are much lower than those of the LMW, and the isotope ratios of the gases are much lower than those of the vapors equilibrated with the LMW at $100\text{ }^\circ\text{C}$.

3.4. Residual gas composition

Figure 5 shows the compositional data for the fumarolic gases at site J plotted on a N_2 -He-Ar ternary diagram. The data

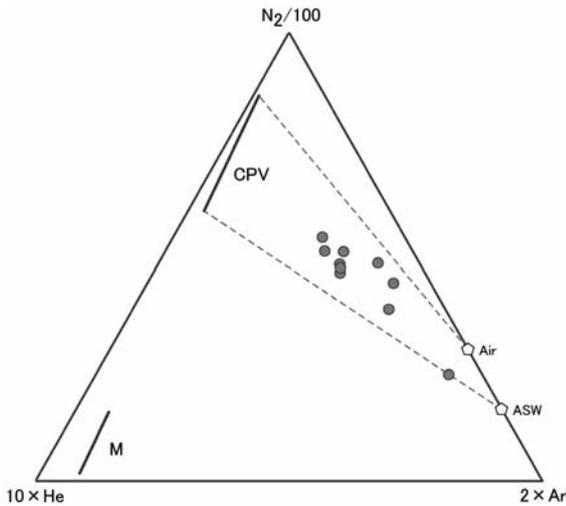


Figure 5. N_2 -He-Ar ternary diagram showing the compositions of the fumarolic gases at site J (molar basis). Circum-Pacific volcanoes (CPV) and M, end members typical of the CPV gases and gases originating in the mantle, respectively. ASW, composition of the water saturated with air. The data are plotted from Table 1.

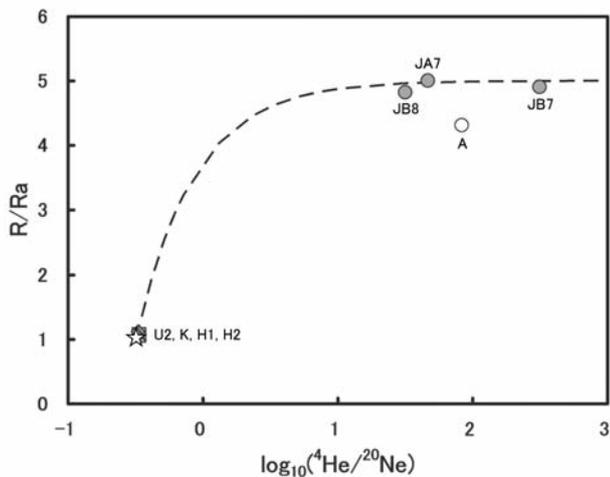


Figure 6. $^3\text{He}/^4\text{He}$ versus $^4\text{He}/^{20}\text{Ne}$ ratios of the fumarolic gases. R and Ra, $^3\text{He}/^4\text{He}$ ratios of fumarolic gas and air (1.4×10^{-6}), respectively. Dashed curve, mixing line between air and an end member with an R/Ra of 5 and a $^4\text{He}/^{20}\text{Ne}$ of 1000.

points are distributed in the region of mixing between the end member with a high N_2/Ar ratio and the atmospheric end members. The end member with a high N_2/Ar ratio can be represented by the gas of the Circum-Pacific volcanoes (Figure 5, CPVs) [Giggenbach 1997]. The $^3\text{He}/^4\text{He}$ ratios in the fumarolic gases are plotted in Figure 6 (JA7, JB7, JB8 and A), where the $^3\text{He}/^4\text{He}$ ratio is normalized by the ratio of air (= R/Ra). The R/Ra and $^4\text{He}/^{20}\text{Ne}$ ratios of U, K and H are almost identical to the atmospheric values. The data points for JA7, JB7 and JB8 plot along the mixing line, the end members of which are air and a component with a high $^4\text{He}/^{20}\text{Ne}$ ratio. The R/Ra ratio of the end member is estimated to have a value of 5, which is slightly lower than the value of 6 reported for He in magmatic gas from the Japanese

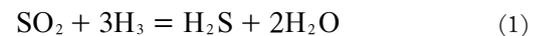
islands [Poreda and Craig 1989]. Fumarolic gas from site A lies beneath the mixing line in Figure 6, suggesting an additional contribution of a ^4He -enriched component that probably originated in the crustal rock.

Figure 7 shows the relationships between the $\delta^{13}\text{C}$ value of CO_2 and the $\text{CO}_2/^3\text{He}$ molar ratios, indicating that the origin of the CO_2 is the fumarolic gases [e.g., Marty et al. 1989, Sano and Marty 1995]. CO_2 in volcanic gases can originate from the upper mantle (e.g., volatiles in the mid-oceanic ridge basalt; Figure 7, MORB), the sea-floor sediment (Figure 7, S), or limestone (Figure 7, L). The CO_2 in fumarolic gases at site J originated mainly from limestone, with a minor contribution of MORB-type CO_2 .

4. Discussion

4.1. Apparent equilibrium temperature

The apparent equilibrium temperature (AET) is an index of volcanic activity [Giggenbach 1987], and it can be regarded as the quenched temperature of volcanic gas. As H_2 gas is the most reactive component in fumarolic gases, and because H_2 included in the following reaction has a stoichiometric coefficient of 3, the AET for the reaction:



is expected to be a proxy of gas temperature in the conduit connecting the gas source and the outlet of the fumarole. The temperature of the gas conduit is initially high, as it is similar to the temperature of the magma at depth, but it decreases with the ascent of the gas. At high temperatures, the equilibrium of Equation (1) shifts to the left side of the reaction. During the ascent of the gas within the conduit, the temperature decreases, resulting in a shift in the

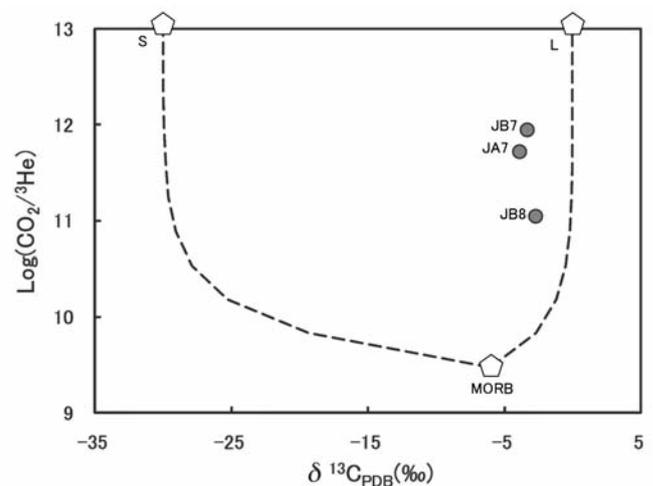


Figure 7. $\text{CO}_2/^3\text{He}$ ratios and $\delta^{13}\text{C}$ values of CO_2 within the fumarolic gases from site J. L, S, end members typical of oceanic sediment (S) and oceanic limestone (L), respectively. Dashed curves, mixing between S and MORB and between L and MORB.

Code	Temp (outlet) °C	AET °C
JA1	135	464
JB1	110	229
JA2	134	263
JB2	106	201
JA3	136	271
JB3	109	273
JA4	143	642
JB4	103	223
JA5	139	482
JB5	98	193
JA6	146	486
JB6	101	240
JA7	111	236
JB7	96	206
JA8	116	379
JB8	97	230

Table 3. Apparent equilibrium temperatures (AETs) for the chemical compositions of JA and JB.

equilibrium of Equation (1) to the right side. However, the reaction takes some time to reach completion, and if the gas ascends rapidly, the shift in the reaction is incomplete, resulting in a high AET. In this case, the increase in the AET is interpreted as reflecting a rapid ascent of the gas.

AET was calculated for the composition of the JA and JB series, as follows:

$$\text{AET}(\text{°C}) = \frac{8924}{\text{LQ} - 1.242} - 273.15 \quad (2)$$

where LQ is the logarithm of the quotient,

$$\text{LQ} = \log \left\{ \left(\frac{C_{\text{H}_2\text{S}}}{C_{\text{H}_2\text{O}}} \right) \left(\frac{C_{\text{SO}_2}}{C_{\text{H}_2\text{O}}} \right)^{-1} \left(\frac{C_{\text{H}_2}}{C_{\text{H}_2\text{O}}} \right)^{-3} \right\} \quad (3)$$

where C_i is the concentration of gas species i . Equation (2) was obtained from the temperature dependence of the standard equilibrium constant of formation, as compiled by Giggenbach [1987]. The total pressure of gas is assumed to be the saturated pressure of water vapor at the calculated AET (see Table 3).

Figure 8 shows the temporal variations in the AET and the outlet temperature of the gas. The AET of the JA series showed a significant increase in 1999, over 1998. The outlet temperature of JA increased slightly in 1999 compared with 1998. Miura et al. [2000] reported that the seismic activity beneath Onawashiro crater was high until September 1998, and decreased thereafter. The increase in AET for the JA series lags behind the period of active seismicity. The AET values for the JB series show no clear temporal trends. The

short distance between fumaroles JA and JB indicates that the processes governing the temporal variations in the AET of JA was active at shallow depths.

4.2. Magmatic vapor at site J

The H_2O isotope ratios of the JA and JB series fall along a line connecting the point for the LMW and the water vapor with a magmatic origin (Figure 4, MV). The δD and $\delta^{18}\text{O}$ values of the MV are similar to those for the JA and JB series; consequently, the line connecting the LMW and the MV passes through the cluster of data points near the MV. The $\delta^{18}\text{O}$ value of the MV was tentatively assumed to be +7‰ relative to the SMOW, which is in the range of typical values for andesitic volcanoes above a subduction zone [Giggenbach 1992]. The $\delta^{18}\text{O}$ value of the MV is limited to less than +8.5‰, because the upper limit for δD of water in andesitic magma is -10‰ [Giggenbach 1992]. The $\delta^{18}\text{O}$ value of JA3 is +5.4‰, close to that of the MV.

JA3 contains high concentrations of HCl. If liquid water encounters a gas similar to that of sample JA3, most of the HCl in the gas would be absorbed by the liquid phase, making the vapor phase depleted in HCl. During September 1998 at Iwate volcano, there was relatively little contamination by water, and consequently there was a higher HCl content in the gas phase. After contamination, no liquid phase remains because it absorbs HCl in the gas phase. For the above case, the temperature of the MV was estimated according to the following equations:

$$\delta_{\text{F}} = \delta_{\text{MV}}f + \delta_{\text{LMW}}(1 - f) \quad (4)$$

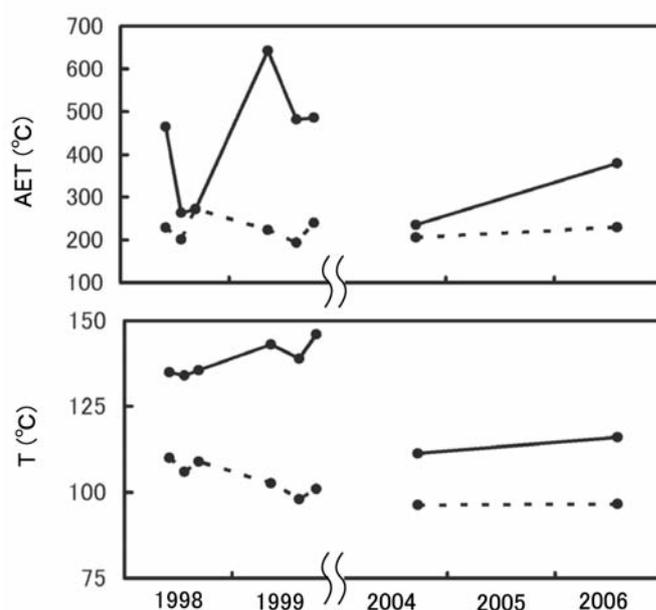


Figure 8. Temporal variations in the apparent equilibrium temperature (AET) calculated for the chemical compositions of the JA (solid lines) and JB (dashed lines) series (upper), and for the outlet temperatures of the gases (lower).

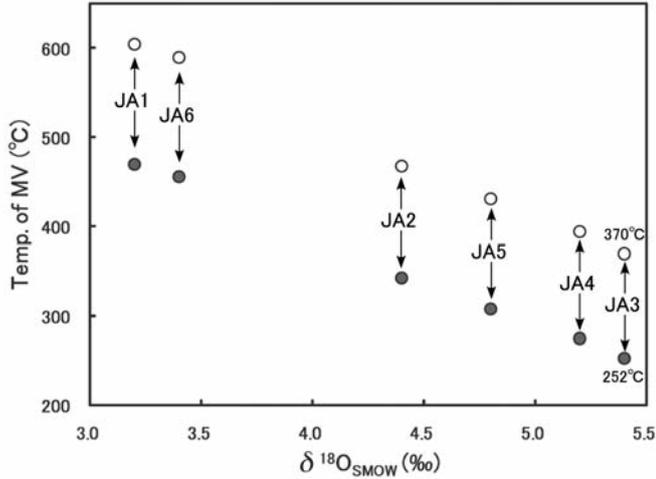


Figure 9. Estimated temperatures of the MV. Filled, open circles, estimates assuming $\delta^{18}\text{O}$ for magmatic vapor of +7‰ and +8.5‰, respectively.

$$H_F = H_{MV}f + H_{LMW}(1 - f) \quad (5)$$

where δ is the $\delta^{18}\text{O}$ value of the gas and the water; the subscripts F, MV and LMW indicate the fumarolic gas, MV and LMW, respectively; f is the molar fraction of H_2O derived from the MV; H is the enthalpy of gas and water; δ_{MV} is assumed to be +7‰ or +8.5‰; δ_{LMW} is assumed to be -11.5‰; and δ_F gives f according to Equation (4). The f obtained is substituted into Equation (5). HF was assumed to be the enthalpy of saturated H_2O vapor at the temperature of the fumarolic gas. H_{LMW} was 42 kJ/kg, which is the enthalpy of liquid water at 10 °C. Solving Equation (5), HMV can be obtained. The temperature of the MV was estimated assuming that the MV is a super-heated water vapor under a pressure of 1 bar (the gas pressure beneath the surface is expected to exceed 1 bar). The temperature of the MV obtained is the potential temperature at the surface when the volcanic gas expands, with the enthalpy of water vapor kept constant. The temperature of the MV depends on δ_F and on the temperature of the fumarolic gas, although the effects of temperature are small relative to those of δ_F .

The calculated temperature is shown in Figure 9. The temperature is lowest for JA3, for which the δ_F value is highest. As discussed below, the JA series is affected by the partial condensation of H_2O vapor, resulting in reduced δ_F values. The calculated temperature of the MV at JA3 is 252 °C and 370 °C for the δ_{MV} values of +7‰ and +8.5‰, respectively. These temperatures are much lower than the general temperatures of degassing magma. It is therefore necessary to invoke a cooling process during the ascent of the MV from the source to the surface.

4.3. Correlation between $\delta^{18}\text{O}$ values of H_2O and $\text{CO}_2/\text{H}_2\text{O}$

The correlation between the chemical and isotopic compositions of fumarolic gases can be explained by a

simple model [Ohwada et al. 2003, Ohba 2007, Ohba et al. 2010] in the case of the Tatun volcanoes (Taiwan), the Izu-Oshima volcano (Japan) and the Kusatsu-Shirane volcano (Japan), for which the fumarolic gas is the vapor phase that is generated between high-temperature magmatic gas and low-temperature groundwater after mixing. As shown in the following equations, a simple mixing between fluids is the basis of the calculations of the composition of the vapor phase, assuming the conservation of total enthalpy with the amount of substance, and the partitioning equilibrium between the vapor and liquid phases.

$$H_{MV}f + H_{LMW}(1 - f) = H_{PS}g + H_{PW}(1 - g) \quad (6)$$

$$\delta_{MV}f + \delta_{LMW}(1 - f) = \delta_{PS}g + \delta_{PW}(1 - g) \quad (7)$$

$$C_{MV}f + C_{LMW}(1 - f) = C_{PS}g + C_{PW}(1 - g) \quad (8)$$

$$\alpha = \frac{\delta_{PW} + 1000}{\delta_{PS} + 1000} \quad (9)$$

$$\beta_{\text{Cl or CO}_2} = \frac{C_{PS}}{C_{PW}} \quad (10)$$

where C is the molar concentration of CO_2 or HCl ; the subscripts PS and PW indicate the vapor phases (primary steam; PS) and the co-existing liquid phase (primary thermal water; PW), respectively, generated from the mixing of the MV and the LMW; and f and g denote the mixing fraction of the MV and the generating fraction of the PS, respectively. In general, g is not equal to f .

Equations (6), (7) and (8) describe the conservation of enthalpy, the isotope ratio, and the chemical composition ($\text{CO}_2/\text{H}_2\text{O}$ or $\text{Cl}/\text{H}_2\text{O}$ ratio) for the generation of the PS and the PW, respectively. Equations (9) and (10) describe the equilibrium between PS and PW in terms of stable isotopes, CO_2 , or the distribution of HCl . α is the isotope fractionation factor of H_2O [Horita and Wesolowski 1994], and β is the distribution coefficient of CO_2 [Giggenbach 1980]. The partitioning of HCl between the vapor phase and the liquid phase is complicated relative to that of CO_2 , because HCl is ionized into H^+ and Cl^- in the liquid phase.

The calculated composition of PS is shown in Figure 10, along with data points for the fumarolic gases. The X axis in Figure 10 is the scaled fraction (ScF) of CO_2 in the CO_2 - H_2O binary system, defined by:

$$\text{ScF}(\text{CO}_2) = \frac{100C_{\text{CO}_2}}{100C_{\text{CO}_2} + C_{\text{H}_2\text{O}}}. \quad (11)$$

The ScF is convenient in terms of plotting a wide range of $\text{CO}_2/\text{H}_2\text{O}$ ratios on a single calibration axis. Given the degree of freedom in f , the composition of PS is expressed as a curve. Smaller values of f give lower $\delta^{18}\text{O}$ values on the PS curve. The PS curve originates from the mixing curve

between the MV and the LMW, and its location depends on the enthalpy of MV. In Figure 10, the enthalpy of MV was assumed to be equivalent to that of vapor heated to 400 °C at 1 bar. In constructing the PS curve, the CO₂/H₂O ratio of the MV was assumed to be 0.008, which was chosen to explain the composition of the gas at site A. The correlation between δ¹⁸O and CO₂/H₂O in gases at site J is not compatible with the composition of PS. As shown above, the gas at site J with a high δ¹⁸O value could be a simple mixture of the MV and the LMW. The generation of a liquid phase is unlikely, because of the high HCl concentration in the gas at site J.

Figure 10 shows a wide range of CO₂/H₂O ratios. Gases with a high CO₂/H₂O ratio have a lower δ¹⁸O value, which can be explained if the gas undergoes partial condensation of the water vapor, modeled by the Rayleigh process. The composition of the gas phase after the partial condensation of the water vapor is expressed as follows:

$$\delta = (\delta_i + 1000)F^{(\alpha-1)} - 1000 \quad (12)$$

$$r = r_i F^{(\frac{1}{\beta}-1)} \quad (13)$$

where F is the fraction of water vapor remaining after the condensation of the water vapor, r is the CO₂/H₂O ratio of the gas phase (molar ratio), and the subscript i represents the initial value of the gas before condensation. The gases from site J lie around the trend of condensation at 100 °C or 120 °C, except JB8 (Figure 10).

The gas samples from U, K and H lie far from the PS curve in Figure 10. The δ¹⁸O values of the gases are much

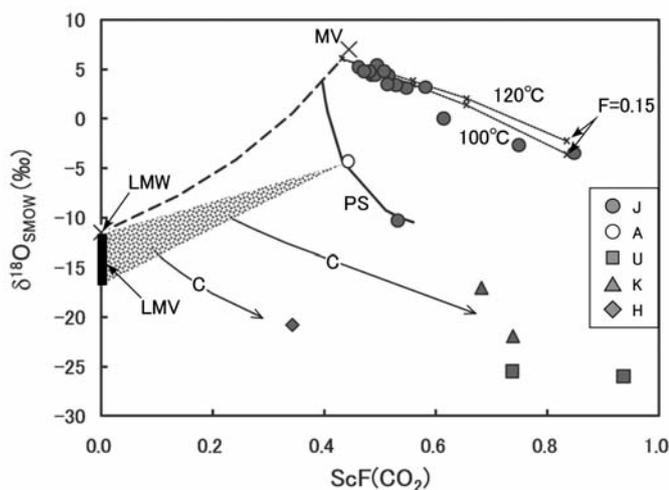


Figure 10. Correlation between the δ¹⁸O values of H₂O and the CO₂/H₂O ratios in the fumarolic gases. Bold dashed curve, mixing line between the MV and the LMW. The LMV represents the vapor generated from the LMW at 100 °C. Shaded region, mixture between the LMV and a gas similar to that from site A. Solid curve (labeled PS), composition of primary steam. Thin dashed curves and solid lines, trends of the partial condensation of the water vapor phase, for which F is the fraction of water vapor remaining after partial condensation.

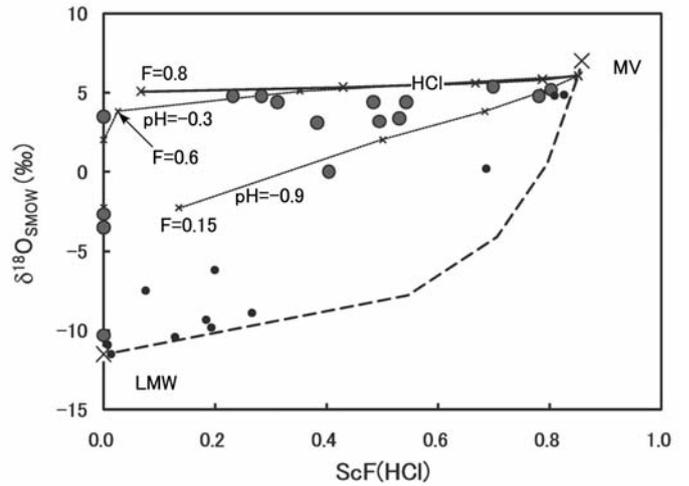


Figure 11. Correlation between the δ¹⁸O values of H₂O and the HCl/H₂O ratios in the fumarolic gases from site J. Bold horizontal line, trend during the partial condensation of the vapor, where the condensed liquid contains H⁺ and Cl⁻ ions with a common molarity. Thin dashed curves, trends of partial condensation, for which the pH of the condensed liquid phase is fixed at -0.3 or -0.9. Bold dashed curve, mixing line between the MV and the LMW.

lower than that of PS. Even if partial condensation is assumed for PS, the δ¹⁸O values would still be too high; however, the δ¹⁸O value can be reduced by introducing a vapor phase that originated from the LMW. If groundwater is heated in the shallow crust within a geothermal area, some of the water would be vaporized. The isotope ratio of the vapor generated (the local meteoric vapor; LMV) is shown on the Y-axis in Figure 10. Given the freedom of the fraction of vaporized water, the δ¹⁸O value of the LMV is expressed as a line. The addition of the LMV to a gas phase (e.g., the gas at A) produces a gas phase with a low CO₂/H₂O ratio, as shown by the shaded region in Figure 10. If such a gas undergoes partial condensation, the composition of the remaining gas phase would approach the composition of the gas from sites U, K and H.

4.4. Correlation between δ¹⁸O values of H₂O and HCl/H₂O

The HCl concentrations in the vapor phase (after partial condensation) can also be calculated under the assumption of a Rayleigh process. Two different cases are assumed in terms of condensation. In the first case, the condensed liquid phase is a pure HCl solution. The concentration of H⁺ in the condensed liquid is equal to that of the Cl⁻ ions. In this case, the HCl concentration in the vapor phase is given by:

$$m = \frac{1}{k} \left(1 - \frac{1}{\sqrt{F}} + \sqrt{\frac{Km_i}{F}} \right)^2 \quad (14)$$

where m is the molarity of HCl in the vapor phase, and K is the partitioning factor of HCl between the vapor and liquid phases, as defined by:

$$K = \frac{m}{m_{\text{H}^+} m_{\text{Cl}^-}} \quad (15)$$

where m_{H^+} and m_{Cl^-} are the molarities of the H^+ and Cl^- ions in the condensed liquid phase, respectively. The temperature dependence of K was given by Simonson and Palmer (1993). The temperature during condensation was assumed to be 120 °C. The changes in the $\delta^{18}\text{O}$ value of H_2O and in the $\text{HCl}/\text{H}_2\text{O}$ ratio of the vapor phase are shown as a sub-horizontal line in Figure 11, in which the X-axis is the ScF of HCl in the $\text{HCl}-\text{H}_2\text{O}$ binary system, which is defined by:

$$\text{ScF}(\text{HCl}) = \frac{1000C_{\text{HCl}}}{1000C_{\text{HCl}} + C_{\text{H}_2\text{O}}} \quad (16)$$

HCl in the vapor phase is effectively removed by the condensed liquid phase. Most of the HCl was lost by the time that 20% of the water vapor had condensed. The above compositional trend does not always explain the distribution of data points for the gases from site J.

We assume the following condition for the second case of condensation. Within a conduit filled with volcanic gas, a condensed liquid phase is likely to be dominant. The liquid phase is likely to be strongly acidic due to the absorption of HCl from the vapor phase; consequently, the removal of HCl from the vapor phase is suppressed because K is the equilibrium constant of the following reaction:



Two cases of condensation were assumed, where the pH of the coexisting liquid phase is -0.3 or -0.9 . When the pH of the condensed liquid phase is fixed, the change in the $\text{HCl}/\text{H}_2\text{O}$ ratio of the vapor phase is given by:

$$m = m_i F^\gamma \quad (18)$$

where,

$$\gamma = \frac{10^{\text{pH}}}{K} - 1. \quad (19)$$

The trend in the composition with progressive condensation is given in Figure 11. Most of the data points for the gases from site J are distributed within the two curves. The composition of the hot spring waters at site J is also shown. Most of the data points for the hot spring water are distributed along the direct mixing curve between the MV and the LMW.

5. Conclusions

Based on the composition of the gases from site J, the AET reached a peak in May 1999. This period with high AET followed the period of intense earthquakes, which suggests that magma movement was followed by magma degassing, which may have increased the flux and speed of

the gas ascent through the conduit to the surface. The temperature of magmatic vapor at site J was estimated at 252 °C to 370 °C, based on the enthalpy and isotope ratio of the gas. During the ascent, contamination by groundwater was limited, which suggests the existence of a well-developed inner wall of the gas conduit, which prevents the intrusion of groundwater. During the ascent of the magmatic gas to the surface, the enthalpy of the gas would have been lost by conduction through the wall of the conduit. The magmatic gas would have ascended a large distance from the deep source to the surface, which is consistent with the findings of a magnetotelluric survey [Aizawa et al. 2009] that revealed an electrically resistive body beneath the geothermal area at site J, at depths of 0.5 km to 3.0 km. The resistive body was interpreted to be a solidified intrusive magma. The gas was probably transported through the solidified magma.

The correlations among the $\delta^{18}\text{O}$ values of H_2O , $\text{CO}_2/\text{H}_2\text{O}$ and $\text{HCl}/\text{H}_2\text{O}$ suggest that the gases at site J experienced partial condensation of the water vapor. During condensation, a strongly acidic liquid phase would have coexisted with the vapor phase, preventing the absorption of HCl in the vapor phase by the liquid phase. The strongly acidic liquid phase might have originated from the condensed liquid phase.

The $\text{CO}_2/\text{H}_2\text{O}$ ratio of the magmatic gas was estimated as 0.008, which is much lower than the 0.028 observed for magmatic gas sampled from the erupting Unzen volcano in 1992 [Ohba et al. 2008]; this suggests that the degassing magma beneath Onawashiro crater is depleted in CO_2 relative to the magma of Unzen volcano.

The fumarolic gas sampled at site A might be a steam phase generated by the mixing of magmatic gas and groundwater. The fumarolic gases at sites U, K and H showed significant condensation of water vapor. The addition of water vapor derived from groundwater would have occurred prior to the condensation.

The N_2 -He-Ar compositions of the gases from site J suggests that the magmatic component is identical to that of andesitic volcanoes above a subduction zone. The correlation between the $^3\text{He}/^4\text{He}$ and $^4\text{He}/^{20}\text{Ne}$ ratios in the fumarolic gases suggests that the gases at site J contain an end member for which the $^3\text{He}/^4\text{He}$ ratio is as high as 5 (R/Ra). The correlation between the $\delta^{13}\text{C}$ value of CO_2 and the $\text{CO}_2/^3\text{He}$ ratio suggests that the origin of the CO_2 in the gases from site J was mainly limestone, with a minor component of MORB-type CO_2 .

The hot spring waters discharged at site J comprise a direct mixture of MV and LMW.

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